ASSESSYOURSELF

Practice. Learn. Succeed

Sub : Science **Grade: 6th** **Date: .....**

Name: ..... **ID No.....**

Time allowed: 1 hour Maximum marks: 25

1. Suggest separation technique which would need to separate the mixture of KCl and NH4 Cl. 1

2. The sea-water can be classified as a homogeneous as well as heterogeneous mixture. Comment. 1

3. Explain why particles of a colloidal solution do not settle when left undisturbed for some time. 1

4. You are given two samples of water labelled as 'A' and 'B'. Sample 'A' boils at 100°C and sample 'B' boils at 105°C. Which sample of water will freeze at 0°C? 1

5. Which of the following are not compounds?

(i) Iodine

(ii) Carbon dioxide

(iii) Carbon

(iv) Chlorine gas. 2

6. On heating, calcium carbonate gets converted into calcium oxide and carbon dioxide.

$$CaCO_3 \xrightarrow{\Delta} CaO + CO_2\uparrow$$

(a) Is this a physical or a chemical change?

(b) Can you prepare one acidic solution by using the products formed in the above process? Write the chemical equations involved. 2

7. Classify each of the following as a physical or a chemical change:

(a) Burning of kerosene in a lantern

(b) Addition of curd to milk

(c) Rising of hot air over a radiator

(d) Cooking of vegetables in a cooker. 2

8. How will you distinguish between two liquids–one pure compound and the other a solution? 2

9. Identify colloids and true solutions from the following:

(a) Fog

(b) Muddy water

(c) Brass

(d) Sugar solution.

10. What is a compound? Differentiate between a compound and a solution. 2

11. A solution contains 37g CuSo4 and 111g of water. Calculate the concentration of mass by mass% of the solution. 3

12. You are given a solution of water and an unknown compound 'X' which boils at 73°C. How will you separate the compound 'X' from water? Draw a labelled diagram. 3